

Name: _____

UNIT 2 OVERVIEW

Topic 2.1 How and why do we study matter?

- Matter and its interactions make up our world.
- Safety is key when working with matter.

Workbook

- Read page 44 (Summary)
- Page 45 (Pure Substances & Mixtures)
- Page 47-48 (Physical and Chemical Properties)
- Page 49 (Safety)
- Practice Quiz: pg 50-54

Topic 2.2 How does the periodic table organize the elements?

- Elements are the building blocks of matter.
- Elements can be organized by their properties.
- The modern periodic table organizes elements in groups and periods.
- Elements are classified as metals, non-metals, or semi-metals.

Workbook

- Read pg 55 (Summary)
- Pg 56-60 (Periodic table)
- Pg 61-63 (Metals, Groups ect)
- Practice Quiz: pg 65-68

Topic 2.3 How can atomic theory explain patterns in the periodic table?

- The structure of atoms can be represented using simple diagrams.
- Elements in chemical groups have similar electron arrangements.
- The periodic table shows how properties of elements change in predictable ways

Workbook

- Read Pg 69-70 (Summary)
- Pg 71-72 (Parts of an Atom)
- Pg 73-74 (Bohr & Valence)
- Pg 75-77 (Trends)
- Practice Quiz: Pg 78-80

Topic 2.4 How do elements combine to form compounds

- Compounds account for the huge variety of matter on Earth.
- Ionic compounds are made of ions.
- Covalent compounds are made of molecules.
- Covalent bonding also occurs in elements and network solids.

Workbook

- Read Pg 81 (Summary)
- Pg 82-86 (Conductivity)
- Pg 87-89 (Covalent Bonding)
- Practice Quiz: Pg 90-92

Topic 2.5 How do we name and write formulas for compounds?

- The chemical name of an ionic compound communicates its composition.
- You can determine the formula of an ionic compound from its name.
- Multivalent metals form more than one ion.
- Polyatomic ions are made up of more than one atom.
- Names and formulas of covalent compounds reflect their molecular structure.

Workbook

- Read Pg 93 (Summary)
- Pg 94-95 (Ionic Compound Naming)
- Pg 96 (Multivalent Metals)
- Pg 97 (Polyatomic Ions)
- Pg 98 (Covalent)
- Practice Quiz: Pg 101-103